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## PATENT SPECIFICATION

692,478

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### COMPLETE SPECIFICATION.

#### A Process for the Treatment of Keratinous Materials and Production of Articles from the Treated Materials.

We, WOLSEY LIMITED, a British Company, of 31 King Street, Leicester, Leicestershire, do hereby declare the invention, for which we pray that a patent may be granted to us, and the method by which it is to be performed, to be particularly described in and by the following statement:—

This invention relates to the treatment of keratin and keratin-containing materials in order to render them soluble in dilute alkali or other solvents which do not chemically degrade the protein. These materials are insoluble in water and all solutions or solvents which do not attack them chemically. Materials falling into this class are the protein constituent of wool and all animal hairs, bristles, horns and epidermal layers such as animal hooves. There are many ways of bringing these into solution (for example, refluxing for many hours with strong mineral acids; treatment with caustic alkali solutions) but in all these the protein is broken down into very small molecular fragments (for example individual amino acids) and the product thus obtained bears no resemblance to the original starting material.

One way of rendering these materials soluble without severe degradation is to treat them with alkaline reducing agents which break the disulphide bond so as to yield sulphhydryl groups. Sodium sulphide and compounds containing the thiol group such as the salts of thioglycollic acid have been used for this purpose. These processes suffer from several disadvantages:—

(1) The reagents used have an objectionable odour which is retained to some extent by the product;

(2) They must be used under severely alkaline conditions which degrade the proteins in general and the attack is not confined to the disulphide bond alone. The product is

therefore discoloured and less suitable for many purposes such as making regenerated protein fibres;—

(3) If less alkaline conditions are used all the protein cannot be brought into solution.

We have now found according to this invention that solutions of peracetic acid, performic acid, perpropionic acid and perbutyric acids oxidise the —S—S— bond to sulphhydryl and sulphonic acid groups without substantial attack of any other part of the protein molecule and in particular without causing any substantial degradation of the main peptide chain. The protein thus oxidised is readily soluble in dilute alkali (for example N/100 ammonia) and solubility is complete except for a small residue which varies from 0% to 5% by weight, depending upon the protein concerned. We have also found that other hydrotropic substances (as hereinafter defined) may be used to dissolve the oxidised product wholly or in part.

Thus, according to the process of the present invention keratin or a keratin-containing material is oxidised with an aqueous solution of a saturated per-aliphatic acid having not more than four carbon atoms in the molecule, at a temperature below 100° C., and the whole or a part of the oxidised protein is dissolved in a solvent which is a hydrotropic substance (as hereinafter defined), preferably a dilute alkali, for example a dilute ammonia solution.

To aid solution of the oxidised protein, wetting agents such as long chain fatty sulphates may often with advantage be added.

The term "hydrotropic substances" as used herein refers to dilute alkalis, e.g. ammonium hydroxide, solutions of cuprammonium compounds; copper ethylenedi-

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amine, phenol containing water, resorcinol containing water, phosphoric acid, and their aqueous solutions, aqueous solutions of sodium oleate, very concentrated aqueous solutions of lithium halides and lithium thiocyanate, concentrated aqueous solutions of zinc chloride, concentrated aqueous solutions of urea and substituted urea derivatives such as guanidines. The protein which dissolves in these substances (heating in some cases being necessary) can be precipitated in most cases on dilution and in all cases by being extruded into strong solution of strong electrolytes, e.g. a strong aqueous salt solution such as sodium sulphate or sodium chloride or into an acidic bath such as N/100 hydrochloric acid, or into alcohols or ketones dissolved in water.

Thus, for example, when oxidised keratin is dissolved in a 20% aqueous urea solution of pH = 8, the oxidised keratin does not precipitate by mere dilution with water, but the addition of a trace of acid is necessary to lower the pH of the diluting solution to 4, or the diluting solution must contain strong electrolyte, e.g. a 20% sodium sulphate solution. From a 20% aqueous urea solution at pH 4, however, the material precipitates as soon as it has been diluted five times with water.

This material is of high molecular weight and may be used as a foodstuff or for making plastic articles or fibres. Thus, the precipitate may be re-dissolved and extruded through a spinnerette into an acid coagulating bath.

If the precipitation of the material is effected by acid then ammonium or other alkali is necessary to redissolve it, but if it is precipitated by a solution of a strong electrolyte or a water-miscible alcohol or ketone, then it will dissolve in water.

The material can be rendered insoluble in dilute alkali by cross linking with well-known reagents such as formaldehyde or heavy metal salts. In general, the soluble material can be insolubilised by most of the processes known in the art of making regenerated protein fibres.

Instead of precipitating the oxidised protein from solution a solid can also be obtained by allowing the solution to evaporate.

Thus, the material is soluble in liquid ammonia to some extent, and can be recovered by allowing the ammonia to evaporate, e.g. if it is desired to obtain fibres by extruding the liquid ammonia solution through a spinnerette into air.

From ammonia, good films can be obtained by casting a film and allowing the ammonia to evaporate. Good fibres can be obtained by spinning the solution into a vacuum or into hot air where the aqueous ammonia is quickly evaporated. The material thus obtained has very similar properties to that obtained by precipitating from the ammonia

solution with acids or strong salt solutions (i.e. the product is again readily soluble in dilute alkali and has a structure which gives an X-ray diffraction diagram corresponding to the material known as  $\alpha$  keratin). This same product is obtained when the material is precipitated with water, strong salt solutions or acid from any of the solvents mentioned above.

When the material which has been dissolved in ammonia and then re-precipitated is extruded from a volatile non-aqueous solvent containing no water or only very little water, such as formic acid, on evaporation of the solvent a material is obtained which is characterised by being soluble much less readily in dilute alkali such as N/10 ammonia and which gives an X-ray diffraction pattern corresponding to the material known as  $\beta$  keratin. Shaped articles obtained in this way in the  $\beta$  form are less elastic but much stronger than those in the  $\alpha$  form. The  $\beta$  materials can be hardened and completely insolubilised with formaldehyde and heavy metal salts in the same way as all protein products.

We have found that wool though of complex structure has an interior consisting of keratin in the  $\alpha$  form but coated with a cuticle which consists largely of the  $\beta$  configuration. We have succeeded in imitating this effect by spinning a solution consisting of 15% of the oxidised wool as obtained with peracetic acid dissolved in N/5 ammonia through a spinnerette into a coagulating bath consisting of 20% sodium sulphate. The fibres were dried and then passed rapidly through formic acid, which was then evaporated. The whole fibre was then hardened and insolubilised with formaldehyde. It is believed that the product obtained consists of a fibre with a highly elastic central portion surrounded by a tough and relatively inelastic outer layer.

Immediately after peracetic acid treatment the keratin is soluble in dilute alkali, e.g. N/10 ammonia, and the other hydrotropic substances, e.g. 100% lithium bromide (100 gms. of salt in 100 ml. of water), 15% urea, 50% zinc chloride, but not in formic acid. However, when the oxidised wool is dissolved in aqueous alkali, urea or lithium salts and then precipitated by dilution with water or dilute acid or strong salt solutions such as 20% sodium sulphate the precipitate obtained is then soluble in formic acid and liquid ammonia. The best method of obtaining a solution in formic acid is thus the following: treat the keratin with a peracid for the requisite time, as described in the Examples below, dissolve the product in N/5 ammonia, precipitate oxidised keratin by adding acid until the pH is less than 4, filter off and dry the precipitate, which then readily dissolves in formic acid in the cold.

The material which is not precipitated by acid is of low molecular weight and can be recovered by evaporating the solution or by salting out. We prefer to use peracetic acid as its aqueous solutions are stable for many hours, whereas those of performic lose oxidising power on standing, and perpropionic acid and the perbutyric acids are more expensive.

In general, when using peracetic acid, a solution of the peracetic acid and acetic acid mixture obtained by reaction between hydrogen peroxide and acetic acid is used direct. The solution need not be acid in reaction and may be neutralised with alkalis (for example caustic soda). As soon as the solution becomes alkaline, however, it becomes very unstable and rapidly loses oxidising power. Even at a pH between 4 and 7 the solution becomes progressively less stable. It may be desirable to work at a pH between 4 and 7, however, as the oxidation reaction proceeds more rapidly in this range than in the more acid solution.

Similar considerations apply to the other peracids employed in the present invention which may also be used in the presence of the corresponding aliphatic acid of the general formula  $C_nH_{2n+1}COOH$ .

To render the protein completely soluble all the cystine must be oxidised and it is preferable to carry on the reaction till more than 80% of the cystine has reacted. The cystine content may be estimated by standard analytical methods. The time for this reaction depends upon (1) the pH as already mentioned; (2) the concentration of peracid which may be varied from the strongest obtainable by reacting the organic acid with hydrogen peroxide without subsequent concentration, which is about 45% by weight if the method of preparation of F. P. Greenspan (Ind. Eng. Chem., 1947; 39, 847) is followed, to very dilute solutions of less than 1% by weight; (3) the temperature of reaction which should, however, be kept below 100° C., as there is severe decomposition of peracid at this temperature and usually we prefer to use 20 to 60° C.; (4) the particle size or fibre diameter of the initial material.

The following Table I shows the solubility in 8N ammonia of the oxidised product obtained by treating wool with peracetic acid for different percentages of cystine oxidised.

TABLE I.	
% Cystine oxidised.	% Soluble in 8N ammonia.
20	8
45	9
75	11
90	91
100	92

The following Examples illustrate how the process of the invention may be carried into effect:

1. 5 gms. of botany wool was suspended with stirring in an aqueous solution containing 45% by weight of peracetic acid for 1 hour at 22° C. After this it was filtered off and washed well about 5.1 gms. of wool being left. (weight increase due to reaction  $-S-S- \rightarrow 2(-SO_3H)$  and  $2(-SO_3^-)$ ). This was then suspended with stirring in 200 ml. of N/100 ammonia to which additions of ammonia were made, as this was used up. After about 1 hour all the wool had dissolved giving a viscous solution and leaving only a fine cloudiness of insoluble material. This was filtered off and found to be about 0.2 gm.

The ammonia solution may be used as such for preparing shaped articles, etc., or it may be acidified with N/10 HCl until all the precipitate has formed. This is collected and found to be 3.5 gms. of a white powder readily soluble in dilute alkali.

2. Example 1 was repeated but the oxidation was carried out for 24 hours in a 1.6% solution of peracetic acid obtained by diluting with water a mixture prepared according to Greenspan's process approximately 25 times. The yield of soluble protein was 70% as in Example 1.

If the reaction with peracetic acid is carried out for 30 mins. less than 50% of the wool is soluble in dilute ammonia.

3. Perpropionic acid was prepared by condensing 1 gm.-mol of 90% hydrogen peroxide with 1.5 gm.-mols of propionic acid in the presence of sulphuric acid and allowing the mixture to stand for 24 hours. 40 ml. of this solution was mixed with 80 ml. of water and 14 gms. of botany wool was suspended therein for 2 hours at room temperature. The wool was then removed and it was found that there was a slight gain due to oxidation. The thus treated wool was then dissolved in 1 litre of 0.2N ammonia and 90% of it went into solution.

The protein can be recovered from this solution by acidification at a pH of less than 5 or by the addition of strong salts, such as 16% by weight of sodium sulphate plus 20% by weight of magnesium sulphate, but other strong electrolytes or mixtures of strong electrolytes work equally as a precipitating medium.

The yield is 60-70%; that is, of each 100 parts by weight of wool which went into solution, 60-70 parts by weight are recovered on acidification or adding strong salts.

4. *n*-Perbutyric acid was prepared by condensing 1 gm.-mol. of 90% hydrogen peroxide with 1.5 gm.-mols. of *n*-butyric acid in the presence of sulphuric acid and allowing the mixture to stand for 24 hours. Example 3 was then repeated using this perbutyric acid instead of perpropionic acid, all the other conditions being the same. Similar results were obtained.

5. From a solution of solubilised  $\alpha$  keratin a regenerated protein fibre was obtained in the following way. The protein was prepared by oxidising solvent extracted virgin merino wool with an aqueous solution containing 16% of peracetic acid for 2 hrs. at room temperature, dissolving the wool in 0.3/N ammonium hydroxide, filtering off the insoluble residue and acidifying to a pH < 4 with 0.1/N hydrochloric acid when a precipitate was formed which was filtered, washed with distilled water and dried. The powder was ground in a mill with 0.3/N ammonium hydroxide to give a solution containing 15% of protein which gelled at room temperature but gave a free flowing viscous solution suitable for spinning at 70° C. Accordingly, the protein solution was extruded at 70° C. through a spinnerette into a coagulating bath at the same temperature consisting of an aqueous solution of N sodium sulphate, 2N magnesium sulphate and 0.1% of the material sold under the Registered Trade Mark "Fixanol" (this is a material manufactured by Imperial Chemical Industries, Ltd. and containing cetyl pyridinium bromide).

White filaments which could be readily handled were thus obtained. The fibres, however, readily dissolved in water and were rendered insoluble by treatment for 30 mins. at room temperature with 2% formaldehyde in a solution of the same composition as the coagulating bath.

The X-ray diagram of the fibre thus obtained has a typical  $\alpha$  structure with a ring at 6.2 Å and a doublet at 10.5 and 9.5 Å, but shows little orientation although some arcing of the reflection of the 10 Å spacings can be detected. After stretching the fibres 100% in water a highly oriented  $\alpha$  diagram is obtained which is almost identical with that of untreated wool. The only difference between the diffraction patterns of the native and regenerated fibre is the presence of weak  $\beta$  reflections in the latter which may be produced either by a small quantity of  $\beta$  protein or by some  $\alpha$ - $\beta$  transformation during the stretching.

This shows that it is possible to solubilise wool and bring it into true solution without losing the  $\alpha$  configuration of the molecule and that on regeneration it crystallises into the same form as when produced by a biological

synthesis. The solubilisation process, although involving extensive chemical reaction does not change the specific  $\alpha$  molecular orientation of the molecule and can thus not be considered as denaturation since this always gives protein having a  $\beta$  structure.

6. 14 gms. of wool (weighed at normal regain) were allowed to stand in 40 ml. of peracetic acid solution (containing 40% of peracetic acid and 60% of acetic acid) and 60 ml. of water for 2 hours at room temperature. The wool was then washed with cold water and stirred with 1 litre of 0.2N ammonium hydroxide. After 2 hours the wool disintegrated and the solution was filtered. The clear yellow filtrate was made just acid to litmus with 2N sulphuric acid. The white curdy precipitate represented 58% of the original dry weight of the wool. The residue was re-treated with peracetic acid and ammonia as described above and the insoluble product obtained was 8.5 to 9.0% of the original dry weight of the wool.

7. Example 6 was repeated but using one twenty-fifth of the concentration of peracetic acid and a time of treatment of 25 hours, and the same result was obtained.

8. 7 gms. of fine cow's horn shavings were stirred with 40 ml. of peracetic acid solution (containing 40% of peracetic acid and 60% of acetic acid) and 60 ml. of water for 2 hours at room temperature. The horn was then washed with cold water and stirred with 1 litre of 0.2 N ammonium hydroxide. After 2 hours the horn disintegrated and the solution was filtered. The filtrate was made just acid to litmus with 2N sulphuric acid. The white curdy precipitate represented 51.5% of the original horn.

9. Example 6 was repeated using propionic acid instead of peracetic acid. The precipitate obtained from the ammonium hydroxide represented 49.5% of the original dry weight of the wool.

10. Example 6 was repeated using *n*-butyric acid instead of peracetic acid. The precipitate obtained from the ammonium hydroxide represented 39% of the original dry weight of the wool.

11. Example 6 was repeated using the same proportions of ingredients but employing peracetic acid solutions of varying pH values. The results are shown in Table II below.

TABLE II.

	pH of peracetic acid solution	% precipitated from ammonium hydroxide soln.	Residue %	% precipitated on acidifying initial soln.
110	2	60	9	0
	4	54	12.7	0
	7	38-52	3.4	3
115	9	17-38	49	0
	10	rapid decomposition		

12. Example 6 was repeated but using aqueous urea saturated at 18° C. and adjusted to pH 8 with sodium bicarbonate instead of ammonium hydroxide for effecting solution.
- 5 The residue was 29% and on acidifying the urea solution with 2N sulphuric acid a white precipitate was obtained amounting to 45.8% of the original dry weight of the wool.
- 10 13. Example 6 was repeated ten times using the following as solvents instead of the ammonium hydroxide:
- (a) 5% aqueous sodium bicarbonate solution (with heating);
  - (b) 100% phosphoric acid;
  - 15 (c) 70% aqueous phosphoric acid (with heating);
  - (d) 30% aqueous sodium oleate (with heating);
  - (e) saturated aqueous zinc chloride solution (with heating);
  - 20 (f) cuprammonium hydroxide;
  - (g) copper ethylenediamine;
  - (h) phenol saturated with water at 18° C. (with heating);
  - 25 (i) 100% lithium bromide (100 gms. of salt in 100 ml. of water) in N/10 hydrochloric acid (with heating);
  - (j) resorcinol containing water (with heating).
- 30 What we claim is:—
1. A process for the treatment of keratin or keratin-containing materials wherein keratin or a keratin-containing material is oxidised with an aqueous solution of a saturated per-aliphatic acid having not more than four carbon atoms in the molecule, at a temperature below 100° C., and the whole or a part of the oxidised protein is dissolved in a solvent which is a hydrotropic substance (as hereinbefore defined).
  - 35 2. A process for the treatment of keratin or keratin-containing materials wherein keratin or a keratin-containing material is oxidised with an aqueous solution of peracetic acid at a temperature below 100° C. and the whole or a part of the oxidised protein is dissolved in a solvent which is a hydrotropic substance (as hereinbefore defined).
  - 40 3. A process as claimed in Claim 1 or 2, wherein the hydrotropic substance is a dilute aqueous alkali, e.g. a dilute ammonia solution.
  4. A process as claimed in any one of the preceding claims, wherein the keratin-containing material is wool.
  - 45 5. A process as claimed in any one of the preceding Claims 1 to 3, wherein the initial material is horn.
  6. A process as claimed in any one of the preceding claims wherein the solution obtained, is treated with a dilute aqueous acid to precipitate a keratinous material.
  7. A process as claimed in any one of the preceding Claims 1 to 5, wherein the solution obtained is salted out with a strong aqueous salt solution to precipitate a keratinous material.
  8. A process as claimed in any one of the preceding Claims 1 to 5, wherein the solution obtained, if necessary after the addition of water, is treated with a water-miscible alcohol or ketone to precipitate a keratinous material.
  9. A process as claimed in any one of the preceding Claims 1 to 5, wherein an aqueous ammonia solution of the oxidised protein is spun or extruded to produce fibres.
  10. A process as claimed in any one of the preceding Claims 1 to 5, wherein an aqueous ammonia solution of the oxidised protein is evaporated in thin layers to produce films.
  - 80 11. A process for the treatment of keratin or keratin-containing materials substantially as described with reference to any one of the Examples given.
  12. Keratinous materials or fibres or films when produced by the process claimed in any one of the preceding claims.
  - 85

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Agents for the Applicants.

PROVISIONAL SPECIFICATION.  
No. 23,858, A.D. 1949.

**A Process for the Treatment of Keratinous Materials and  
Production of Articles from the Treated Materials.**

We, WOLSEY LIMITED, a British Company, of 31, King Street, Leicester, Leicestershire, do hereby declare the nature of this invention to be as follows:—

This invention relates to the treatment of keratin and other protein materials of high cystine content in order to render them water-soluble. These materials are insoluble in

water and all solutions or solvents which do not attack them chemically. Materials falling into this class are the protein constituent of wool and all animal hairs, bristles, feathers, horns and epidermal layers such as animal hooves. There are many ways of bringing these into solution (for example refluxing for many hours with strong mineral

acids; treatment with caustic alkali solutions) but in all these the protein is broken down into very small molecular fragments (for example individual amino acids) and the product thus obtained bears no resemblance to the original starting material.

One way of rendering these materials soluble without severe degradation is to treat them with alkaline reducing agents which break the disulphide bond so as to yield sulphhydryl groups. Sodium sulphide and compounds containing the thiol group such as the salts of thioglycolic acid have been used for this purpose. These processes suffer from several disadvantages:

(1) The reagents used have an objectionable odour which is retained to some extent by the product;

(2) They must be used under severely alkaline conditions which degrade the proteins in general and the attack is not confined to the disulphide bond alone. The product is therefore discoloured and less suitable for many purposes such as making regenerated protein fibres;

(3) If less alkaline conditions are used all the protein cannot be brought into solution.

We have now found according to this invention that solutions of peracetic acid and performic acid oxidise the  $-S-S-$  bond to  $-SO_3H$  without attacking appreciably any other part of the protein molecule. The protein thus oxidised is readily soluble in dilute alkali (for example N/100 ammonia) and solubility is complete except for a small residue which varies from 0% to 5% by weight, depending upon the protein concerned.

Thus, according to the process of the present invention keratin or related protein material of high cystine content is oxidised with a solution of peracetic acid or performic acid and the oxidised protein is dissolved in a dilute alkali, for example a dilute ammonia solution.

To aid solution of the oxidised protein, wetting agents such as long chain fatty sulphonates may often with advantage be added. The protein may be recovered by acidifying the solution and at a pH < 7 from 40% to 80% by weight of the original material precipitates out. If the alkaline solution used to dissolve the protein contains a high concentration of urea (for example > 20% by weight of urea) it requires much lower pH values for precipitation and this may for some purposes be advantageous. This material is of high molecular weight and may be used as a foodstuff or for making plastic articles or fibres. Thus, the precipitate may be redissolved and extruded through a spinnerette into an acid coagulating bath. The material can be rendered insoluble in dilute alkali by cross linking with well known reagents such as formaldehyde or heavy

metal salts. In general, the soluble material can be insolubilised by most of the processes known in the art of making regenerated protein fibres.

The material which is not precipitated by acid is of low molecular weight and can be recovered by evaporating the solution or by salting out. The oxidation with the peracids is carried out from aqueous solution. We prefer to use peracetic acid as its aqueous solutions are stable for many hours, whereas those of performic lose oxidising power on standing.

In general, a solution of the peracetic acid and acetic acid mixture obtained by reaction between hydrogen peroxide and acetic acid is used direct. The solution need not be acid in reaction and may be neutralised with alkalis (for example caustic soda). As soon as the solution becomes alkaline, however, it becomes very unstable and rapidly loses oxidising power. Even at a pH between 4 and 7 the solution becomes progressively less stable. It may be desirable to work at a pH between 4 and 7, however, as the oxidation reaction proceeds more rapidly than in the acid solution.

To render the protein completely soluble all the cystine must be oxidised and it is preferable to carry on the reaction till more than 90% of the cystine has reacted. The cystine content may be estimated by standard analytical methods. The time for this reaction depends upon (1) the pH as already mentioned; (2) the concentration of peracid which may be varied from the strongest obtainable, which is about 45% by weight if the method of preparation of F. P. Greenspan (Ind. Eng. Chem., 1947, 39, 847) is followed, to very dilute solutions of less than 1% by weight; (3) the temperature of reaction which should, however, be kept below 100° C., as there is severe decomposition of peracid at this temperature and usually we prefer to use 20 to 60%.

The following examples illustrate how the process of the invention may be carried into effect:

1. 5 gms. of botany wool is suspended with stirring in a solution containing 45% by weight of peracetic acid for 1 hour at 22° C. After this it is filtered off and washed well, about 5.4 gms. of wool being left (weight increase due to reaction  $-S-S- \rightarrow 2(-SO_3H)$ ). This is then suspended with stirring in 200 c.c.s. of N/100 ammonia to which additions of ammonia are made, as this gets used up. After about 1 hour all the wool has dissolved giving a viscous solution and leaving only a fine cloudiness of insoluble material. This is filtered off and found to be about 0.2 gm.

The ammonia solution may be used as such for preparing shaped articles, etc., or it may be acidified with N/10 HCl until all the pre-

cipitate has formed. This is collected and found to be 3.5 gms. of a white powder readily soluble in dilute alkali.

2. Example 1 is repeated but the oxidation is carried out for 24 hours in a 1% solution of peracetic acid. The yield of soluble protein is 70% as in Example 1.

If the reaction with peracetic acid is carried out for 30 mins. less than 50% of the wool is soluble in dilute ammonia.

3. Example 1 is repeated but finely powdered bullocks horn is used instead of

wool. All the oxidised horn is soluble in ammonia and there is no insoluble fraction.

Dated the 15th day of September, 1949.

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Consulting Chemists and  
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Bank Chambers, 329, High Holborn,  
London, W.C.1,  
Agents for the Applicants.

# PROVISIONAL SPECIFICATION.

No. 4582, A.D. 1950.

## A. Process for the Treatment of Keratinous Materials and Production of Articles from the Treated Materials.

15 We, WOLSEY LIMITED, a British Company, of 31, King Street, Leicester, Leicestershire, do hereby declare this invention to be described in the following statement:—

20 This invention relates to the treatment of keratin and other protein materials of high cystine content in order to render them water-soluble and is concerned with a development of the invention of our Specification No. 23858/49.

25 The said Specification No. 23858/49 describes a process wherein keratin or related protein material of high cystine content is oxidised with a solution of peracetic acid or performic acid and the oxidised protein is dissolved in a dilute alkali, for example a dilute ammonia solution.

30 It has now been found according to this invention that the process of Specification No. 23858/49 can be carried out using perpropionic acid or perbutyric acid instead of peracetic acid or performic acid.

35 Thus, according to the process of the present invention, keratin or related protein material of high cystine content is oxidised with a solution of perpropionic acid or perbutyric acid and the oxidised protein is dissolved in a dilute alkali, for example, a dilute ammonia solution.

40 The following examples illustrate how the process of the invention may be carried into effect:

45 1. Perpropionic acid was prepared by condensing 1 gm.-mol. of hydrogen peroxide with 1.5 gm.-mols of propionic acid in the presence of sulphuric acid and allowing the mixture to stand for 24 hours. 40 c.cs. of

this solution was mixed with 60 c.cs. of water and 14 gms. of Botany wool was suspended therein for 2 hours at room temperature. The wool was then removed and it was found that there was a slight gain due to oxidation. The thus treated wool was then dissolved in 1 litre of 0.2N ammonia and 90% of it went into solution.

50 The protein can be recovered from this solution by acidification at a pH of less than 6 or by the addition of strong salts, such as 16% by weight of sodium sulphate plus 20% by weight of magnesium sulphate, but other strong electrolytes or mixtures of strong electrolytes work equally well as a precipitating medium.

55 The yield is 60—70%; that is, of each 100 parts by weight of wool which went into solution, 60—70 parts by weight are recovered on acidification or adding strong salts.

2. Perbutyric acid was prepared by condensing 1 gm.-mol. of hydrogen peroxide with 1.5 gm.-mols. of butyric acid in the presence of sulphuric acid and allowing the mixture to stand for 24 hours. Example 1 was then repeated using perbutyric acid instead of perpropionic acid, all the other conditions being the same. Similar results were obtained.

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## PROVISIONAL SPECIFICATION.

No. 8897, A.D. 1950.

A Process for the Treatment of Keratinous Materials and  
Production of Articles from the Treated Materials.

We, WOLSEY LIMITED, a British Company, of 31, King Street, Leicester, do hereby declare this invention to be described in the following statement:—

5 This invention relates to the manufacture of shaped articles from keratin and other protein materials of high cystine content.

10 These materials, which will hereinafter be referred to collectively as "keratin" are insoluble in water and all solutions or solvents which do not attack them chemically. Materials falling into this class are the protein constituents of wool and all animal hairs, bristles, feathers, horns and epidermal layers such as animal hooves.

15 Specifications Nos. 23858/49 and 4582/50 describe processes for treating keratin wherein the keratin is oxidised by means of performic acid, peracetic acid, perpropionic acid and perbutyric acid, and the solution of the oxidised keratin is dissolved in dilute alkali, for example a dilute ammonia solution from which the oxidised keratin may be re-precipitated by the addition of dilute acid and thereafter re-dissolved in dilute alkali.

20 A solution in alkali of the oxidised keratin can be used for the preparation of shaped articles such as fibres or sheets which can be obtained by extruding the alkaline solution into an acid medium or by casting from an ammonia solution and allowing the ammonia solution to evaporate. The products thus obtained can be rendered insoluble by treatment with formaldehyde, various metal salts or other processes known for insolubilising and strengthening products obtained from proteins such as casein or vegetable globulins.

25 We have now discovered according to this invention that the oxidised keratin obtained from peracids can be precipitated from solution not only by acids but by alcohols and ketones dissolved in water as well as by solutions of strong electrolytes, e.g. 20% sodium sulphate.

30 We have also discovered that a number of solvents other than dilute alkali may be used to facilitate the solution of the oxidised keratin and that the use of these solutions facilitates the production of shaped articles by extrusion of these solutions.

35 The solvents are all characterised by being able to break hydrogen bonds and are known in chemistry as hydrotropic substances; they include solutions of cuprammonium; phenols, and their aqueous solutions, very concentrated aqueous solutions of lithium

halides and lithium thiocyanate, concentrated aqueous solutions of zinc chloride, concentrated aqueous solutions of urea and substituted urea derivatives such as guanidine. From these solutions the protein is precipitated in most cases on dilution and in all cases by being extruded into strong solution of simple electrolytes such as sodium sulphate or sodium chloride or into an acidic bath such as N/100 hydrochloric acid.

40 Thus, for example, while oxidised keratin is dissolved in a 20% aqueous urea solution of pH = 8, the oxidised keratin does not precipitate by mere dilution with water, but the addition of a trace of acid is necessary to lower the pH of the diluting solution to 4, or the diluting solution must contain strong electrolyte, e.g. a 20% sodium sulphate solution. From a 20% aqueous urea solution at pH 4, however, the material precipitates as soon as it has been diluted five times with water. The material is also soluble in liquid ammonia and can be recovered by allowing the ammonia to evaporate, e.g. if it is desired to obtain fibres by extruding the liquid ammonia solution through a spinnerette into air.

45 Instead of precipitating the oxidised protein from solution a solid can also be obtained by allowing the solution to evaporate. From dilute aqueous alkali solutions, in particular ammonia, good films can be obtained by casting a film and allowing the ammonia to evaporate. Good fibres can be obtained by spinning the solution into a vacuum or into hot air where the aqueous ammonia is quickly evaporated. The material thus obtained has very similar properties to that obtained by precipitating from the ammonia solution with acids or strong salt solutions (i.e. the product is again readily soluble in dilute alkali and has a structure which gives an X-ray diffraction diagram corresponding to the material known as  $\alpha$  100 keratin). This same product is obtained when the material is precipitated with water, strong salt solutions or acid from any of the solvents mentioned above.

50 When the material is extruded from a volatile nonaqueous solvent containing no water or only very little water, such as formic acid, or liquid ammonia, on evaporation of the solvent a material is obtained which is characterised by being soluble much less readily in dilute alkali such as N/10 ammonia and which gives an X-ray diffraction pattern



corresponding to the material known as  $\beta$  keratin. Shaped articles obtained in this way in the  $\beta$  form are less elastic but much stronger than those in the  $\alpha$  form. The  $\beta$  materials can be hardened and completely insolubilised with formaldehyde and heavy metal salts in the same way as all protein products.

We have found that wool though of complex structure has an interior consisting of keratin in the  $\alpha$  form but coated with a cuticle which consists largely of the  $\beta$  configuration. We have succeeded in imitating this effect by spinning a solution consisting of 15% of the oxidised wool as obtained with peracetic acid dissolved in N/5 ammonia through a spinnerette into a coagulating bath consisting of 20% sodium sulphate. The fibres were dried and then passed for five seconds through formic acid and then passed into a vacuum in which the formic acid evaporated. The whole fibre was then hardened and insolubilised with formaldehyde. The product obtained consisted of a fibre with a highly elastic central portion surrounded by a tough and relatively inelastic outer layer and closely resembled a natural wool fibre.

Immediately after peracetic acid treatment the keratin is soluble in dilute alkali, e.g.

N/10 ammonia, and some of the solvents mentioned, e.g. 100% lithium bromide (100 gms. of salt in 100 c.c.s. of water), 15% urea, 50% zinc chloride, but not in others such as formic acid and liquid ammonia and soluble only with difficulty in phenol. However, when the oxidised wool is dissolved in alkali, urea or lithium salts and then precipitated by dilution with water or dilute acid or strong salt solutions such as 20% sodium sulphate the precipitate obtained is then soluble in formic acid and liquid ammonia. The best method of obtaining a solution in formic acid is thus the following:— treat the keratin with a peracid for the requisite time as described in Specification No. 23858/49 or 4582/50, dissolve the product in N/5 ammonia, precipitate oxidised keratin by adding acid until the pH is less than 4, filter off and dry the precipitate, which then readily dissolves in formic acid in the cold.

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